

“Did You Know That Perfectly Balanced Electrolytes Produce A Charge In Your Body, Regulating Metabolic Functions You Need For Superior Athletic Performance and Training and Helping Decrease The Chances Of BURN OUT?”

Electrolytes are the most overlooked aspect of sports nutrition. Many endurance athletes have a misconception about how to replace electrolytes effectively, yet it's the one KEY component to athletic nutrition!



Have You Been In A Race Or A Workout And Felt God Awful?

“Burn out” can easily occur when your body is deficient in key vitamins and minerals. REAL Electrolytes are composed of Sodium, Potassium, Calcium, Magnesium, Chloride, Bicarbonate, Phosphorus, and Sulfur. Individually these electrolytes cannot provide the conductivity necessary to transmit the electrical impulses that operate our nerves, muscles and heart. Together, in ratio, they regulate the nervous system and muscle functions, regulate the amount of water in body fluids and cells, and balance the body's alkaline/acid ratio which is essential to fighting disease.



Electrolytes give the body the electrical charge that can have you training harder. Diet, climate, stress, illness, medications and exercise all impact electrolyte levels. Climate such as cold weather or hot weather can upset electrolyte balance and cause dehydration.

Warning: Just taking these minerals individually WILL NOT give you a balance replacement.

Taking Salt Does Not Cure An Electrolyte Imbalance

Although salt IS an electrolyte component, it's just one of the 7 components that need replacement in the body. There are problems with taking too much or too little salt.

Elevated levels of sodium create electrolyte imbalances which have been linked to high blood pressure and heart problems. The impact of low sodium receives almost no attention. Low levels of sodium can lead to low blood pressure and create susceptibility to bacteria such as staph and E.coli.

....And that's just **one** example of **one** component of electrolytes being imbalanced!

Salt levels are just one of the components that must be balanced for superior athletic function, as well as maintaining overall health.

Electrolytes Cannot Be Replaced With Water or Sports Drinks

The biggest misconception about electrolyte imbalances and treating dehydration is that you can drink lots of water or a sports drink. Plain water will dilute electrolytes, creating a greater electrolyte imbalance.

"Sports" drinks are missing several essential electrolytes and are used mostly for quick carb replacement. The sugar in sports drinks also kills the natural electrical charge from a real electrolyte solution. This charge is critical for absorption into the cells.

The Truth About “Electrolyte” Products & Replacement of Electrolytes:



- **Mineral “fizzy” packets** do not contain proper ratios of electrolytes and are missing most of the vital components of a real electrolyte solution.
- **Sports drinks and gels** are for carb replacement. They may contain a little potassium or sodium but this still won't replace and balance electrolytes.
- **Mineral water** can further eschew the balance of electrolytes contained in your body.
- **Drinking water** increases the dilution of electrolytes, increasing low levels of vital electrolytes.
- **Athletes can die or get seriously ill** from an improper balance of electrolytes or low levels of electrolytes. Low blood pressure combined with low heart rate is a very dangerous thing. Electrolytes can actually normalize your blood pressure which will have you racing and training faster.
- **Electrolyte “mixes”** that you add to water, are not the real thing. They cannot balance your electrolytes. They won't create an electrical charge
- It's **critical for endurance athletes** to replace and balance electrolytes. Most of the population is electrolyte deficient which can cause many types of illnesses and diseases. It's important for everyone to have the proper balance.

Why is A Balanced Replacement So Important?

Perfectly balanced electrolytes actually produce an electrical charge! The charge regulates a host of metabolic functions your body needs for training.

Electrolyte replacement is NOT just needed for the summer months. Winter can be deceiving. You think you sweat less, but you can actually lose electrolytes when you're not sweating. Winter has a drying effect and when training or racing, it's imperative to replace electrolytes.

Electrolytes are lost through the intestinal tract, sweat and urination. Diet, drinking habits, medication, activity level, climate, stress and illness all impact electrolyte levels.

How Do You Know If You Need Additional Electrolytes?

- **Sweat excessively or don't sweat easily**
- **Have cold hands and feet**
- **Have low or high blood pressure**
- **Have diet habits**
- **Take medications**
- **Have existing illness or disease**
- **Exercise regularly**

Many People Are Affected By Electrolyte Imbalances

There is considerable evidence that more and more people are affected by electrolyte imbalances. Heat related illnesses and deaths are on the rise. Every year athletes collapse and sometimes die...many resulting from electrolyte imbalances.

Athletes and the elderly aren't the only ones affected by electrolyte imbalances. Millions take prescriptions, such as diuretics, which alter electrolyte levels leading to additional health concerns. Other medications that affect electrolyte balances are: antihistamines, phenothiazines (tranquilizers), tricyclic antidepressants, beta blockers and vasoconstrictors.

General loss of electrolytes occur from: running, gardening, mowing, farming, playing golf, walking, driving and many other activities! So everyone is susceptible to having an electrolyte imbalance.

But if you're an athlete, it's absolutely critical to replace vital electrolytes. Those seeking to be highly competitive, can get an edge on the competition by replacing and balancing electrolytes!

How Do You Get Electrolytes?

Various food and fluids contain electrolyte sources. While getting electrolytes through diet is an essential component for good health, it's nearly impossible to get all of the essential electrolytes in the proper ratios. Historically people replaced electrolytes with bone broths and soups. Today's diet furnishes too little or too much of individual electrolytes. With an electrolyte imbalance causing so many health ailments, the replacement of electrolytes should be a priority for everyone.

Until now, there has been NO real balanced electrolyte drink replacement available....

“Who Else Wants To Improve Their Athletic Performance By Taking The ONLY Milliequivalent, Isotonic Electrolyte Drink On The Market Today?”

This balanced formula provides just the right amount of perfectly balanced electrolytes in a milliequivalent formula that is super absorbable, without adding unnecessary carbohydrate calories. It contains NOTHING artificial and is processed so carefully that this isotonic formula is the same concentration as your blood. The potency and balance is so strong that opening the bottle sometimes produces an electrical shock similar to static shock.



The late renowned chemist who developed this highly regarded solution has passed the process to his family and has been offering this product to the public quietly over the past 10 years. 35,000 studies were conducted on the effectiveness. The patented solution is kept safe and away from large companies who want to buy and ultimately dilute the potency of the product.

What's The Best Way To Replace Vital Electrolytes and Keep Them Balanced?

The best way to replace vital electrolytes and keep them balanced is by taking Life Balances Electrolytes. They have the only formula on the market that is completely balanced, isotonic and charged.

"This electrolyte replacement is made from a proprietary process with a melt-emulsion reaction in a potentized ionic formula with a free floating orbital. "Inversion Vortices" is the process for development of electrolytes in solution which is based on gustatory (taste) pathways. The biochemical individual can determine how much electrolyte fluid replacement is needed at that time."

In laymen's terms this means that Life Balances Electrolytes is the ONLY really balanced electrolyte drink available. These electrolytes are perfectly balanced and will not upset the body's mineral balance or add unnecessary carbohydrate calories.

One Of The Most Unique Features Is That It's Taste Sensitive!

This patented solution will actually tell you BY TASTE if you've replaced enough electrolytes or not. By following the simple directions you can insure you're balanced optimally on a daily basis.

And because everyone's body chemistry is different, some athletes will lose more than others during workouts and racing.

Now there is a way that each individual athlete can taste their way to a perfect balance! That's truly unique!

How Do I Use Life Balances Electrolytes?

The first thing you should know is that these electrolytes come in a concentrated form, which means they can be easily taken on trips and used at meets and competitions.

These electrolytes are offered in three formulas: Blue Label, Red Label and Green Label.

The formula is a milliequivalent liquid formula which is ESSENTIAL to balancing electrolytes.

Which Formula Is Right For Me?

Everyone can use the Green Label formula. The other formulas are more specialized for people who know they have high or low levels of sodium and potassium.

Green Label

This contains normal ratios of electrolytes and is best suited for people who have normal sodium and potassium levels and have normal blood pressure.

Blue Label

This contains a higher level of sodium and is appropriate for those who have low levels of sodium and may have low blood pressure. People who use this formula also have low levels of calcium and magnesium.

Red Label

This contains higher levels of potassium and is appropriate for people with low levels of potassium, phosphorus and nitrogen, and may have high blood pressure. People who use this formula sometimes take diuretics.

If you don't know your blood pressure or blood chemistry levels, use the **Green Label**.

What Does The Electrolyte Replacement Taste Like?

This is a serious product. It's medical grade, which means there are no sweeteners or flavorings added. The mixture, however, tastes good to most people simply because most people are very deficient in electrolytes. The taste and texture actually changes from person to person depending on how deficient their levels are.

At the beginning the solution to a very deficient person will taste thick and salty. If it tastes really bad, capillary circulation is poor. By taking the electrolyte this will change over time because the formula will increase circulation and once levels are restored, the taste will be mild and slightly salty.

This electrolyte is so special that it's taste sensitive and each person can achieve a near perfect balance of electrolytes by using their own taste buds!

Get Ahead Of Your Competition By Replacing and Balancing Electrolytes- A Fundamental, But Overlooked Task When It Comes To Athletic Nutrition!

Make 2009 Your BEST Athletic Year Ever!



We want you to run fast, bike fast, swim fast and train harder!

Order Today

3 Ways To Order



Balanced Electrolyte Replacement Guaranteed!

"35,000 studies can't be wrong! We're THAT SURE you'll love this amazing electrolyte formula. If you're not satisfied, just return the unused portion and we'll give you a full refund -no hassles and no hard feelings!"



Yes! I would love to have the athletic edge. Send me my Electrolyte Solution Kit which makes 4 gallons of highly potent, electrically charged electrolyte. My order has a 30-day risk-free guarantee! I will receive one complete Electrolyte Solution Kit for only \$53.95 plus \$7.95 shipping & handling.

One 8 oz concentrate



=

1 Gallon of Electrolyte



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The following articles are written by John Kitkoski, creator of the Taste Sensitive Electrolytes:

The Science Behind TRUE Electrolytes *It's not just for athletes!*

Whether you are speaking about plants, animals, or humans, electrolytes are essential for the existence of life. They provide the means, when combined with proper circulation and balanced body chemistry, to maintain proper blood pressure, circulate nutrients, and rebuild damaged tissues and expedite waste from the body, including our lymphatic circulatory system.

Electrolytes in the body are minerals such as sodium, potassium, chloride etc. that are dissolved in the blood. When the electrolytes are dissolved they break apart into charged particles called ions. The ions carry either a negative or positive charge. These charged particles create electricity (much like a common battery) that help run the bodies of animals and humans.

Electrolytes are the basis of good health because they are used in the maintenance and repair of all tissue, the utilization of amino acids (cells and tissue “building blocks”), and as the basis of every physical and neurological function. They maintain osmotic equilibrium of our cell walls and the internal water balance that enables muscles and nerves to contract and expand, and wounds to heal. They are also essential for growth and development of the bones as well as the organs.

Electrolytes are responsible for carrying minerals and amino acids to all points of the body. The pH (the measure of acidity or alkalinity) of the body is regulated by the electrolytes. Without electrolyte balance (homeostasis) the body may be in a disease state, nutrient deficient and vulnerable to disease.

Electrolytes help to chelate toxic and inorganic minerals out of the body before they cause damage, and they also assist in the absorption of minerals into the cells where the body can use them.

Electrolytes are the primary factors for cell and immune system health. They keep cell membranes strong, raise their osmotic pressure, so that no virus or bacteria can enter, and maintain correct pH in intracellular fluid so that invading organisms cannot survive.

Healthy cells are kept round by osmotic pressure or osmotic equilibrium – an opposing fluid pressure between the inside of the cell and on the outside of the cell. It is a critical balance and protects the autoimmune system.

It is time for scientist to realize that there is something deeper in man that has to be cared for. We cannot heal a person through symptomatic relief. We must go to the root of the problem – to the cell itself – where electrolytes are at working giving life. It is here that real healing can begin.

Electrolytes supply the spark to cells. They are not fuel that is burned to provide power. Rather, like a spark in a car engine, they are the electric fire that ignites every chemical reaction in a cell. They deliver electrons where needed for reactions, and store charges between events.

Too much positive charge from acids (+) creates an inability to circulate electrons. Excess anions (-) of an alkaline state will overcharge a cell or organism.

Life is a balancing act, so most biological processes need neutral – or slightly alkaline pH to assure a steady supply of electrons. Our blood remains very close to 7.45 pH, and if it changes by .1, we can die. Electrolytes not only help restore neutral pH balance, they also act as buffers that resist any change in pH.

A “neutral factor” “means if we eat something that is too acid or alkaline, our body can prevent a change in pH. A healthy body with electrolytes will temporarily neutralize these extremes. Electrolytes that contain bicarbonates, create neutral factor to buffer body fluids and neutralize anions and cations if positive or negative charge is too strong.”

A cell’s most critical chore is to maintain the integrity of the cell membrane – the inner-outer pressure balance at a cell wall to separate cell from not-cell. This double layer of lipids is in constant motion, fed by electrolytes. A strong membrane is a cell’s first line of defense – the front-line of the immune system. Without electrolytes, this barrier can not be sustained, weakens, and a virus or bacteria can invade the cell.

Electrolytes are key catalysts in thousands of enzymes needed by the cells to make amino acids, proteins, and other organic molecules. When electrolytes form, they generate more electro-chemical activity, attract more minerals, capture more charge. Charge control is the key to enzymes that allow biochemical reactions to occur rapidly, selectively, precisely. Zinc, for one, is used in over 20 enzyme systems.

For example, the pancreas produces enzymes and acids to break food down in digestion. Drinking electrolytes 30 minutes before a meal moistens and recharges soft tissues lining the digestive tract. Then, when you eat, membranes and micro-organisms are ready to digest and absorb food. But further, electrolytes supply the pancreas new ions to make more digestive enzymes and hydrochloric acid.

Electrolytes are also critical to nerves-both individually, and collective coordination of the entire nervous system. Nerve impulses are transmitted as an exchange of sodium and potassium ions at the nerve membrane and through the nerve synapse centers. Nerve membrane is encased in long tendrils of protein with a calcium ion attached at the end of each strand. Without this impulse of ion fire, there can be no taste, no smell, no sight, no sensation, no awareness.

Hormones, vitamins and enzymes which activate, regulate and synchronize nerve action all require a mineral ion as key element in their reactive structure, and for their synthesis. For one, cobalt is needed by the pineal gland to make melatonin, the hormone which regulates neurologic function to determine the level of sleep or wakefulness.

This touches only a few of the many profound, essential roles of electrolytes in blood chemistry, cell biology, human physiology, and brain psychology. Electrolytes are the key to unlock energy flow in a cell. They strike the sparks of electric fire that make life happen.

FLUID REPLACEMENT (ELECTROLYTES) AND THE Biochemical Individual

Old-time American ranchers used to drink batches of **switchel** – a mixture of water, molasses, and vinegar – during haying season. Workouts have changed since then, and, thankfully, so have the energy drinks. The water is still there, but the molasses and vinegar have been replaced with some form of sugar, minerals such as potassium and sodium, and ever-mysterious “natural flavors.” **The end result is basically the same: a double shot of water and high fructose corn syrup.**

Most people believe a balance of 1 sodium and 2 potassium ratio can help the intake of water when in sports drinks but if you get enough of them in your diet they may not be any help for the first couple hours of a work out.

During prolonged exercise in the heat, people can become dehydrated at a rate of 1-2 L every hour (about 2-4 lbs of body weight loss per hour). Even a slight amount of dehydration causes physiological consequences. For example, every liter (2.2 lbs) of water lost will cause heart rate to be elevated by about eight beats per minute, cardiac output to decline by 1 L/min, and core temperature to rise by 0.3° C when an individual participates in prolonged exercise in the heat. Sweating is the way in which the body maintains its' core temperature at 37 degrees centigrade. This results in the loss of body fluid and electrolytes and if unchecked will lead to dehydration and eventually circulatory collapse and heat stroke. To minimize disturbances in cardiovascular function and body temperature and to reduce the perceived difficulty of exercise, people should attempt to drink fluids at close to the same rate that they are losing body water by sweating.

The decision as to how much fluid to ingest should be based upon a risk-benefit analysis. Undoubtedly, the most serious consequence of inadequate fluid replacement, i.e., dehydration, hyperthermia, which when severe will cause heat exhaustion, heat stroke, and even death. The risks of too much fluid ingestion are gastrointestinal discomfort and physical difficulty of drinking large volumes of fluid.

There are three types of drinks all of which contain various levels of fluid and electrolytes: Isotonic, Hypotonic, and Hypertonic. Isotonic quickly replaces fluids lost by sweating and supplies a boost of carbohydrate. Glucose is the body's preferred source of energy.

Hypotonic quickly replaces fluids lost by sweating, suitable for those who need fluid without the boost of carbohydrates.

Hypertonic use is to supplement daily carbohydrate intake to top up muscle glycogen stores. If used during exercise Hypertonic drinks need to be used in conjunction with isotonic drinks to replace fluids.

Electrolytes serve three general functions in the body: many are essential minerals, they control osmosis or water between body compartments and they help maintain the acid-base balance required for normal cellular activities. The sweat the evaporates from the skin contains a variety of electrolytes. The electrolyte composition of sweat is variable but comprises of the

following components: Sodium, Potassium, Calcium, Magnesium, Chloride, Bicarbonate, Phosphate and Sulphate.

Lactated Ringer's Solution (LRS) is a balanced (electrolyte concentration similar to serum) and isotonic (osmolality similar to serum) solution. $\text{Na}^+=131$, $\text{K}^+=4$, $\text{Ca}^{++}=3$, $\text{Cl}^-=110$, $\text{Lactate}^-=28\text{mEq/L}$. LRS is non-acidifying; and provides small amounts of K^+ , and large amounts of Na^+ and Cl^- .

Lactate^- combines with H^+ to form lactic acid, which is metabolized to CO_2 and H_2O by the liver. Lactated Ringers, therefore, yields HCO_3^- (or more correctly, consumes H^+) over a period of time as a function of liver metabolism of lactic acid. LRS is an excellent ECF replacement fluid, also of use in Metabolic Acidosis. It is the most commonly used fluid for a multitude of disease processes in all species.

Polyionic R-148 is also balanced and isotonic. $\text{Na}^+=140$, $\text{K}^+=5$, $\text{Mg}^{++}=3$, $\text{Cl}^-=98$, $\text{Acetate}^-=27$, $\text{Gluconate}^-=23\text{ mEq/L}$. Acetate^- and gluconate^- consume H^+ similar to Lactate^- . This solution is interchangeable with LRS. Since acetate is metabolized by more tissues than the liver, it is a better bicarbonate precursor and thus has theoretical advantage.

Polyionic R-148D5 and D5LR are solutions of polyionic R-148, or lactated Ringers, which also have 50 grams of Dextrose added per 1000 ml of solution. They are balanced and hypertonic (approximately 550 mOSm/L). The only real problem with these solutions is that they are about 2 x hypertonic. The possibility of phlebitis is increased. They should be given through a central vein and should not be used if there is already hyperosmolar or hyperglycemic. The addition of glucose as an energy source and to prevent hypoglycemia can be very useful.

The knowledge about the fluids of the body during health and illness is still relatively meager. The composition of these fluids is well established, but the factors that influence their rate of movement are practically unknown. The soda pop companies and the sports drinks know what to add as ingredients to get you to drink more, more, more. But they do not know what to add to make you balanced and well, well, well. The latter are especially important because the rates of exchange of the reactants in the multitude of chemical reactions that proceed simultaneously in the living must be somewhat dependent on the rate of exchange of the fluids of the body.

EQUILIBRIUM

The disturbances in fluid balance may be classified as: (1) disturbances in volume; (2) disturbances in electrolytic osmolar concentration; (3) disturbances in composition; (4) disturbances in distribution; (5) disturbances in the rate of internal exchange.

*(John L. Kitkoski (the created of Taste Sensitive Electrolytes) took the mathematical development which dealt with the kinetics of conduction of electrons in biological solids and applied it to conduction of **ions** in cells because in some respects the cell resembles a solid rather than a liquid.)*

For conventional ion transport, the cell is regarded as a semi-permeable bag containing a solution of ions in liquid water. However, the state of intracellular water has been shown to be

nonliquid, and major fractions of intracellular Na^+ and K^+ have been shown to exist in a complexed state. Therefore, it seems more appropriate to consider the cell as an organized, nonliquid phase, consisting of **macromolecules** embedded in a matrix of crystalline water. Intracellular Na^+ and K^+ ions may then be supposed to locate themselves mostly in complexed form on sites on the macromolecules (analogous to a valence band), but to be capable of mobility by hopping from site to site through the crystalline water in which the Na^+ and K^+ are only sparingly soluble (analogous to a conduction band). Such a picture leads to the concept that the conduction and potential of Na^+ and K^+ in the cell conform to mathematical laws analogous to those governing electrons in semiconductor solids.

Cell water is organized in layer of polarized water molecules, arranged in 10-20 concentric layers around each individual protein molecule. Symmetry of cell water is spherical (or ellipsoidal) around each protein molecule, but with the different spherical arrays randomly positioned relative to each other.

The driving potential for a current across the cell wall of a single species of ion what a single positive charge is the difference in chemical potentials of that type of ion between the inside and outside of the cell. Counter-ions are assumed able to move freely to maintain macroscopic charge neutrality. Interaction of potentials or of currents with other species of ions are assumed not to exist.

The essential chemical difference between the plasma and intercellular compartments relates solely to protein, with concentration of 16 mEq/L and 1 mEq/L per liter, respectively, (For body electrolyte purposes, proteins are considered to be anions.) Otherwise the ionic patterns of the two compartments are so close that they are generally considered to be one in most clinical situations relating to electrolyte balance. The chief extracellular cation is by far Na^+ with an average of 142 mEq/L. The concentration of K^+ and Ca^{++} may each be taken as 5 mEq/L and that of Mg^{++} as 2 mEq/L. Hardly as a surprise, Cl^- , with a concentration of 103 mEq/L, is the chief anion. The average concentrations of HCO_3^- and $\text{HPO}_4^{=}$ are 27 and 2 respectively. The total concentration of cation and total concentration of anion hover, in health, close to the value of 155 mEq/L.

Life Balances electrolytes are a melt-emulsion reaction, potentized ionic formula with a free floating orbital. Based on the Gustatory (taste) pathways an individual can determine how much electrolyte fluid replacement is needed at that time. If an individual is deficient, the electrolyte fluid replacement will taste thick and bad. If an individual has good fluid level, the electrolyte fluid replacement will have a thinner texture and a milder taste. If the individual has excesses, the electrolyte fluid replacement will be thin and bitter tasting.

The Science Behind Our Electrolyte Formula!

There are many factors that affect electrolyte balance. Every second of every day, our bodies rely on electrolytes just to support normal physiological functions, and as a result, we experience small, daily losses of electrolytes. For instance, the average loss of fluids and electrolytes through perspiration can total 600 milliliters (ml) per day. In respiration, this amount is 400 ml; in feces, 200 ml and in urination, 1300 ml.

Other factors such as pregnancy, poor diet, dehydration, use of diuretics, disease, exertion, vomiting, diarrhea, trauma and excessive perspiration significantly increase one's need for electrolytes.

UNDERSTANDING ELECTROLYTES

Electrolyte is a substance whose molecules split into individual ions when dissolved thus allowing it to conduct electrical energies. But we must go back further to understand how to make an electrolyte replacement drink...

Step 1. Minerals: Minerals are naturally occurring inorganic elements having a characteristic crystalline structure and chemical composition. Minerals or macro-minerals are those minerals which the body requires more than 100 milligrams of per day for proper maintenance of health. Macro or trace minerals are those minerals which the body requires more than 100 milligrams of per day.

Some of the most essential minerals and trace minerals have shown a dependency upon being ionic in order to be absorbed in the intestinal lumen and/or to be physiologically valid.

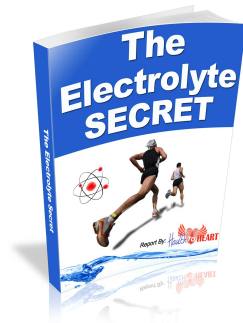
Step 2. Ionic: Of, containing, or relating to an ion or ions. Through digestive processes, some of which are dependent upon sufficient stomach acid, the body is able to break down some other forms of organic or inorganic bound minerals and trace minerals to their ionic form so they can be absorbed. The Life Balances formulation of melt-emulsion breaks down the bound minerals and trace minerals to their ionic form so as to be capable of remaining in free-flowing, non-bound ionic form in a balanced solution.

An ion is an atom or group of atoms that are inherently either positively or negatively charged due to either additional or missing electron(s). This charge causes the ions to interact, attracting or repelling each other in a search for another ion to join with or to give up an electron in order to make the charge neutral.

Inversion vortexing allows for orbitals of atoms to become free floating in a matrix of solution as fluids of the body are largely ionic solutions. The body uses the movement of ions through these fluids and across cell membranes as an intricate part of vital body processes. For example, ions regulate acid-base balance and water balance. Ions also serve essential roles in nerve conduction, muscle contraction, bone and tooth formation, and enzyme activation. In fact, every body process is dependant on ions.

Step:3 Consume Life Balances Electrolytes- The ONLY ionic electrolyte on the market!

Order Now!



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